

## Appendix 6: Suggested practical investigations

The following suggestions are *additional* practical investigations that exemplify the scientific process. They can be used to supplement students' understanding of chemistry in addition to the practical investigations found within the main body of the content.

- Investigate the ease of thermal decomposition of carbonates, including calcium carbonate, zinc carbonate and copper carbonate.
- Compare the temperature rise produced when the same volume of water is heated by different fuels.
- Investigate the volume of air used up and products formed when candles are burned.
- Investigate the reactions of calcium compounds: the decomposition of calcium carbonate and the reaction of calcium oxide with water; the reaction of calcium carbonate with acid.
- Carry out simple neutralisation reactions of acids, using metal oxides, hydroxides and/or carbonates.
- Carry out electrolysis of sea water/acidified water.
- Investigate the rusting of iron.
- Investigate simple oxidation and reduction reactions, such as burning elements in oxygen or competition reactions between metals and metal oxides.
- Investigate the fractional distillation of synthetic crude oil and the ease of ignition and viscosity of the fractions.
- Investigate the products produced from the complete combustion of a hydrocarbon.
- Investigate the cracking of paraffin oil.
- Investigate the properties of a group of elements, e.g. Group 2.
- Investigate the properties of typical ionic compounds.
- Test predictions of whether a precipitate forms when soluble salts are mixed.
- Carry out a series of ion tests to identify unknown compounds.
- Build models of simple covalent molecules.
- Investigate the typical properties of simple and giant covalent compounds.
- Investigate the rate of reactions, such as magnesium and hydrochloric acid; or sodium thiosulfate and hydrochloric acid.
- Determine the formula of a hydrated salt such as barium chloride or copper sulfate by heating to drive off water of crystallisation.
- Prepare a substance and calculate the % yield, given the theoretical yield.
- Evaporate a solution to dryness to determine the mass of solute in a given mass of solution.
- Investigate the mass changes at the electrodes during the electrolysis of copper sulfate solution using copper electrodes.
- Investigate the migration of ions in, e.g. potassium manganate (VII) solution.
- Electroplate a metal object.
- Determine the volume of one mole of hydrogen gas by using the reaction of magnesium with hydrochloric acid.
- Determine the molar volume by measuring the volume and mass of a gas (e.g. carbon dioxide).

- Investigate simple reversible reactions, such as the decomposition of ammonium chloride.

Safety is an overriding requirement for all practical work. Centres are responsible for ensuring that whenever their students complete practical work appropriate safety procedures are followed.